

CHEMICAL REACTIONS AND EQUATIONS

1. Which of the following is a feasible reaction ? 2017
- (1) $\text{Ba(s)} + \text{K}_2\text{SO}_4\text{(aq)} \rightarrow \text{BaSO}_4\text{(aq)} + 2\text{K(s)}$
 - (2) $\text{Zn(s)} + 2\text{AgNO}_3\text{(aq)} \rightarrow \text{Zn(NO}_3)_2\text{(aq)} + 2\text{Ag(s)}$
 - (3) $\text{Mg(s)} + \text{Na}_2\text{SO}_4\text{(aq)} \rightarrow \text{MgSO}_4\text{(aq)} + 2\text{Na(s)}$
 - (4) $\text{Cu(s)} + \text{MgSO}_4\text{(aq)} \rightarrow \text{CuSO}_4\text{(aq)} + \text{Mg(s)}$
2. The metal used to recover copper from an aqueous solution of copper sulphate is 2016
- (1) Na (2) Ag (3) Hg (4) Fe
3. Magnesium ribbon is rubbed with sand paper before making it to burn. The reason of rubbing the ribbon is to
- (1) remove moisture condensed over the surface of ribbon.
 - (2) generate heat due to exothermic reaction.
 - (3) remove magnesium oxide formed over the surface of magnesium.
 - (4) mix silicon from sand paper (silicon dioxide with magnesium for lowering ignition temperature of the ribbon).
4. The reaction that differs from the rest of the reactions given is
- (1) formation of calcium oxide from limestone.
 - (2) formation of aluminium from aluminium oxide.
 - (3) formation of sodium carbonate from sodium hydrogen carbonate.
 - (4) formation of mercury from mercuric oxide.
5. What mass of CO_2 will be formed when 6 g of carbon is burnt in 32 g of oxygen? 2015
- (1) 38 g (2) 12 g (3) 26 g (4) 22 g
6. The reaction between carbon and oxygen can be represented as 2013
- $$\text{C(s)} + \text{O}_2\text{(g)} \longrightarrow \text{CO}_2\text{(g)} + \text{heat}$$
- In which of the following type(s), the above reaction can be classified?
- I. Combustion reaction
 - II. Displacement reaction
 - III. Endothermic reaction
 - IV. Combination reaction

(1) (I) and (III)

(2) (I), (III) and (IV)

(3) (I) and (IV)

(4) (I) only

ANSWER

1.2

2.4

3.3

4.2

5.4

6.3

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